

Name: Ahmed Faizan Date: 12 May Period:

Specific Heat and Heat Capacity Worksheet

DIRECTIONS: Use $q = (m)(C_p)(\Delta T)$ to solve the following problems. Show all work and units.

Ex: How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$\begin{aligned} q &= ? & \Delta t &= 55 - 22 = 33^\circ\text{C} & q &= m s \Delta t \\ m &= 10\text{g} & s &= 0.9 \text{ J/g}^\circ\text{C} & &= 10 \times 0.9 \times 33 \\ & & & & &= \underline{297 \text{ J}} \end{aligned}$$

1. The temperature of 335 g of water changed from 24.5°C to 26.4°C. How much heat did this sample absorb? C for water = 4.18 J/g°C

$$\begin{aligned} \Delta t &= 26.4 - 24.5 = 2 & q &= m s \Delta t \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= 335 \times 4.18 \times 2 \\ m &= 335\text{g} & &= \underline{2800.6 \text{ J}} \end{aligned}$$

2. How much heat in kilojoules has to be removed from 225g of water to lower its temperature from 25.0°C to 10.0°C?

$$\begin{aligned} \Delta t &= 25 - 10 = 15 & q &= m s \Delta t \\ m &= 225\text{g} & &= 225 \times 4.18 \times 15 \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= \underline{14107.5 \text{ J}} \\ & & 1\text{J} &= 1 \times 10^{-3} \text{ KJ} & \therefore & \underline{14.1075 \text{ KJ}} \end{aligned}$$

3. To bring 1.0kg of water from 25°C to 99°C takes how much heat input?

$$\begin{aligned} m &= 1\text{kg} = 1000\text{g} & q &= 1000 \times 4.18 \times 74 \\ \Delta t &= 99 - 25 = 74 & & \\ s &= 4.18 \text{ J/g}^\circ\text{C} & q &= \underline{309320 \text{ J}} \end{aligned}$$

Imp An insulated cup contains 75.0g of water at 24.00°C. A 26.00g sample of metal at 82.25°C is added. The final temperature of the water and metal is 28.34°C. What is the specific heat of the metal?

$$\begin{aligned} q_{\text{water}} &= -q_{\text{metal}} & (C = m s) \rightarrow s &= \frac{C}{m} \\ q_{\text{water}} &= 75 \times 4.18 \times (28.34 - 24) & q &= m s \Delta t \\ &= 1252 & 1252 &= 26 \times s \times (82.25 - 28.34) \\ & & 1252 &= 26 \times 53.91 \times s \\ & & & s = \underline{0.893 \text{ J/g}^\circ\text{C}} \end{aligned}$$

5. A calorimeter has a heat capacity of 1265 J/°C. A reaction causes the temperature of the calorimeter to change from 22.34°C to 25.12°C. How many joules of heat were released in this process?

$$\begin{aligned} C &= 1265 \text{ J/}^\circ\text{C} \\ \Delta t &= 25.12 - 22.34 = 2.78 \\ q &= ? \\ q &= C \Delta t \\ &= 1265 \times 2.78 \\ &= \underline{3516.7 \text{ J}} \end{aligned}$$

Answers Of Specific Heat

Leo J. Malone, Theodore O. Dolter



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Answers to Questions Set at Examinations Great Britain. Education Department. Department of Science and Art, 1889

Answers to questions set at the science and art examinations Science and art department, 1887 Matriculation model

answers: heat and light, magnetism and electricity. Being the papers from June 1888 to Jan. 1891, with answers by H.E.

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Table of Contents Answers Of Specific Heat

1. Understanding the eBook Answers Of Specific Heat
 - The Rise of Digital Reading Answers Of Specific Heat
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Of Specific Heat
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Of Specific Heat
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Of Specific Heat
 - Personalized Recommendations
 - Answers Of Specific Heat User Reviews and Ratings
 - Answers Of Specific Heat and Bestseller Lists
5. Accessing Answers Of Specific Heat Free and Paid eBooks

- Answers Of Specific Heat Public Domain eBooks
 - Answers Of Specific Heat eBook Subscription Services
 - Answers Of Specific Heat Budget-Friendly Options
6. Navigating Answers Of Specific Heat eBook Formats
- ePub, PDF, MOBI, and More
 - Answers Of Specific Heat Compatibility with Devices
 - Answers Of Specific Heat Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Answers Of Specific Heat
 - Highlighting and Note-Taking Answers Of Specific Heat
 - Interactive Elements Answers Of Specific Heat
8. Staying Engaged with Answers Of Specific Heat
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Of Specific Heat
9. Balancing eBooks and Physical Books Answers Of Specific Heat
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Of Specific Heat
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers Of Specific Heat
- Setting Reading Goals Answers Of Specific Heat
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers Of Specific Heat
- Fact-Checking eBook Content of Answers Of Specific Heat
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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